

1. **Amphoteric Oxide** – An oxide that reacts with acids and alkalis to form salts.
2. **Base** – Metal Oxide or Hydroxide
3. **Catalyst** – A substance that increases the rate of a chemical reaction, but remains unchanged at the end of reaction.
4. **Condensation Reaction** – Organic molecules join to produce bigger molecules, losing a small molecule; water
5. **Cracking** – A reaction in which big hydrogen molecules are broken into smaller molecules by heat.
6. **Empirical Formula** – the simplest formula of a compound which shows the ratio between the atoms of each element.
7. **Endothermic** – Take in heat, Take in Energy, To break bonds! (Test tube turns hot)
8. **Exothermic** – Release heat, releases energy, For bond-forming! (Test tube turns cold)
9. **Homologous Series** – A set of organic compounds in which the formula of each one differs from the previous one by an extra  $-CH_2-$  group of atoms.
10. **Macromolecule** – A large molecule containing large no. of atoms joined together.
11. **Monomer** – A small molecule which can join together in large numbers to form on big molecule.
12. **Oxidation state** – The charge on an ion.
13. **Polymerisation** – A reaction in which large numbers of similar small molecules are joined together to form one big molecule (Polymer)
14. **Redox reaction** – A reaction in which oxidation & reduction takes place.
15. **Relative atomic mass** - The mass of an atom of an element compared with  $1/12^{\text{th}}$  of the mass of an atom of carbon-12.
16. **Relative molecular mass** – The mass of a molecule of a substance compared with  $1/12^{\text{th}}$  of the mass of an atom of carbon-12.
17. **Solvent** – A liquid used to dissolve a solid. Eg. Ethanol.